

Table of Common Ions

Cations (+ charge)		Anions (- charge)	
Aluminum	$\text{Al}^{+3}$	Acetate	$\text{CH}_3\text{COO}^{-1}$
Ammonium	$\text{NH}_4^{+1}$	Bromide	$\text{Br}^{-1}$
Antimony	$\text{Sb}^{+3}$	Carbonate	$\text{CO}_3^{-2}$
Arsenic	$\text{As}^{+3}$	Bicarbonate (hydrogen carbonate)	$\text{HCO}_3^{-1}$
Barium	$\text{Ba}^{+2}$	Hypochlorite	$\text{ClO}^{-1}$
Bismuth	$\text{Bi}^{+3}$	Chlorite	$\text{ClO}_2^{-1}$
Calcium	$\text{Ca}^{+2}$	Chlorate	$\text{ClO}_3^{-1}$
Chromium (II) (chromous)	$\text{Cr}^{+2}$	Perchlorate	$\text{ClO}_4^{-1}$
Chromium (III) (chromic)	$\text{Cr}^{+3}$	Chloride	$\text{Cl}^{-1}$
Cobalt	$\text{Co}^{+2}$	Chromate	$\text{CrO}_4^{-2}$
Copper (I) (cuprous)	$\text{Cu}^{+1}$	Dichromate	$\text{Cr}_2\text{O}_7^{-2}$
Copper (II) (cupric)	$\text{Cu}^{+2}$	Fluoride	$\text{F}^{-1}$
Hydrogen	$\text{H}^{+1}$	Hydroxide	$\text{OH}^{-1}$
Hydronium	$\text{H}_3\text{O}^{+1}$	Iodide	$\text{I}^{-1}$
Iron (II) (ferrous)	$\text{Fe}^{+2}$	Nitrate	$\text{NO}_3^{-1}$
Iron (III) (ferric)	$\text{Fe}^{+3}$	Nitrite	$\text{NO}_2^{-1}$
Lead (II) (plumbous)	$\text{Pb}^{+2}$	Oxalate	$\text{C}_2\text{O}_4^{-2}$
Lead (IV) (plumbic)	$\text{Pb}^{+4}$	Hydrogen Oxalate (binoxalate)	$\text{HC}_2\text{O}_4^{-1}$
Lithium	$\text{Li}^{+1}$	Oxide	$\text{O}^{-2}$
Magnesium	$\text{Mg}^{+2}$	Peroxide	$\text{O}_2^{-2}$
Manganese (II)	$\text{Mn}^{+2}$	Permanganate	$\text{MnO}_4^{-1}$
Mercury (I) (mercurous)	$\text{Hg}_2^{+2}$	Phosphate	$\text{PO}_4^{-3}$
Mercury (II) (mercuric)	$\text{Hg}^{+2}$	Monohydrogen Phosphate	$\text{HPO}_4^{-2}$
Nickel	$\text{Ni}^{+2}$	Dihydrogen Phosphate	$\text{H}_2\text{PO}_4^{-1}$
Potassium	$\text{K}^{+1}$	Sulfate	$\text{SO}_4^{-2}$
Silver	$\text{Ag}^{+1}$	Sulfite	$\text{SO}_3^{-2}$
Sodium	$\text{Na}^{+1}$	Hydrogen Sulfate (bisulfate)	$\text{HSO}_4^{-1}$
Strontium	$\text{Sr}^{+2}$	Hydrogen Sulfite (bisulfite)	$\text{HSO}_3^{-1}$
Tin (II) (stannous)	$\text{Sn}^{+2}$	Sulfide	$\text{S}^{-2}$
Tin (IV) (stannic)	$\text{Sn}^{+4}$	Hydrogen Sulfide (bisulfide)	$\text{HS}^{-1}$
Zinc	$\text{Zn}^{+2}$	Thiosulfate	$\text{S}_2\text{O}_3^{-2}$