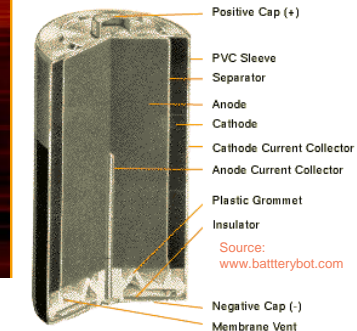
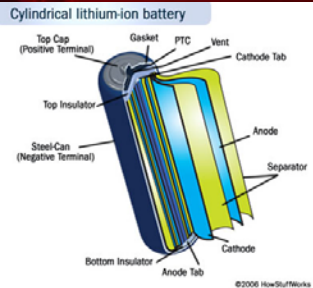
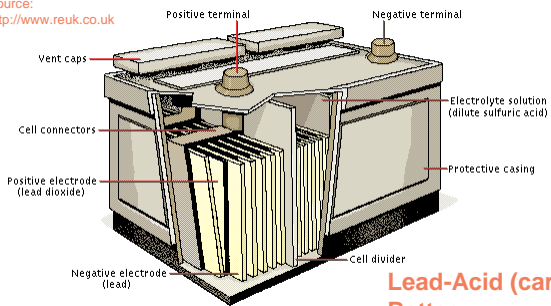


The Electrochemical Cell



Source: <http://www.reuk.co.uk>



Alkaline Battery

The Electrochemical Cell

- The species being oxidized loses electrons and is the reducing agent
 - The species being oxidized (putting out electrons into solution) is the anode
 - The species that is more active (see activity series) will go from metal to ion, thus putting out electrons – therefore the more active metal will be the anode.
 - Remember “anode = oxidized” (both start with vowels) or AN OX | RED CAT (anode oxidized and cathode reduced)

The Electrochemical Cell

- The species being reduced gains electrons and is the oxidizing agent
 - The species being reduced (taking in electrons to make metal out of ion) is the cathode
 - The species that is less active (see activity series) will go from ion to metal, thus taking in electrons – therefore the less active metal will be the cathode.
 - Remember “cathode = reduced” (both start with consonants)

The Electrochemical Cell

- Electrons will flow from anode to cathode (from negative toward positive) through the wire
- Below are diagrams of the Daniell Cell (left from <http://www.chem1.com>)

